

Marcus Hush Theory

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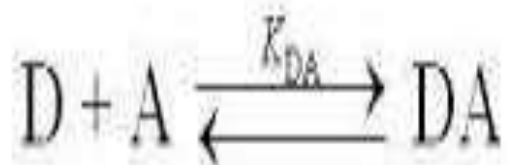
Marcus Hush Theory

Marcus theory for ETR was originally developed by Rudolph A. Marcus in 1956 & due to this theory received Nobel Prize in chemistry in 1992. This Theory is widely used in chemistry as well as biology to explain the rate of ETR. Firstly it was formulated to explain OSETR then Hush extended to ISETR. So it is also known as MH Theory. This theory based upon the TST which gives idea about structural changes in a reaction.

This theory based upon the following consideration:

1. The charge may be transferred in any portion from one body to another.
2. It depends upon different fast electron polarisation P_e and the slow atom and orientation polarisation P_u of the solvent.
3. Marcus separates the inner sphere (reactant + tightly bound solvent molecules, in complexes + ligands) and the outer sphere (free solvent)
4. It must be confined to calculating the outer sphere energy of the non-equilibrium polarization of the "transition state". The outer sphere energy is often much larger than the inner sphere contribution because of the far reaching electrostatic forces.

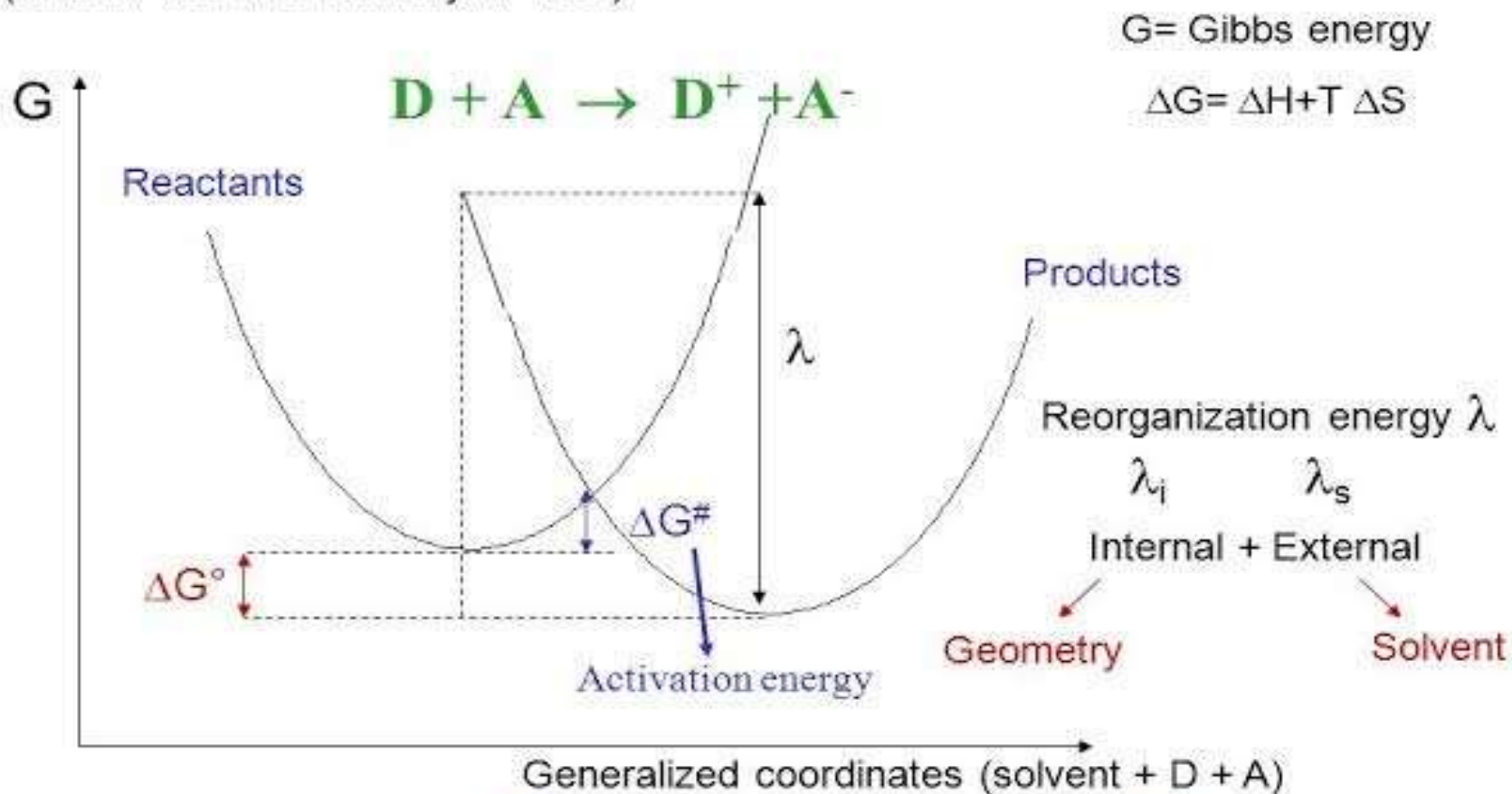
Suppose, as a special case, we have a redox reaction where the electron donor, D, and the acceptor A, reversibly form an encounter complex; where the rate determining step is the electron transfer step; and where break up of the successor complex is fast



This theory mathematically & graphically developed by following way

Marcus theory: ET from a donor (D) to an acceptor (A)

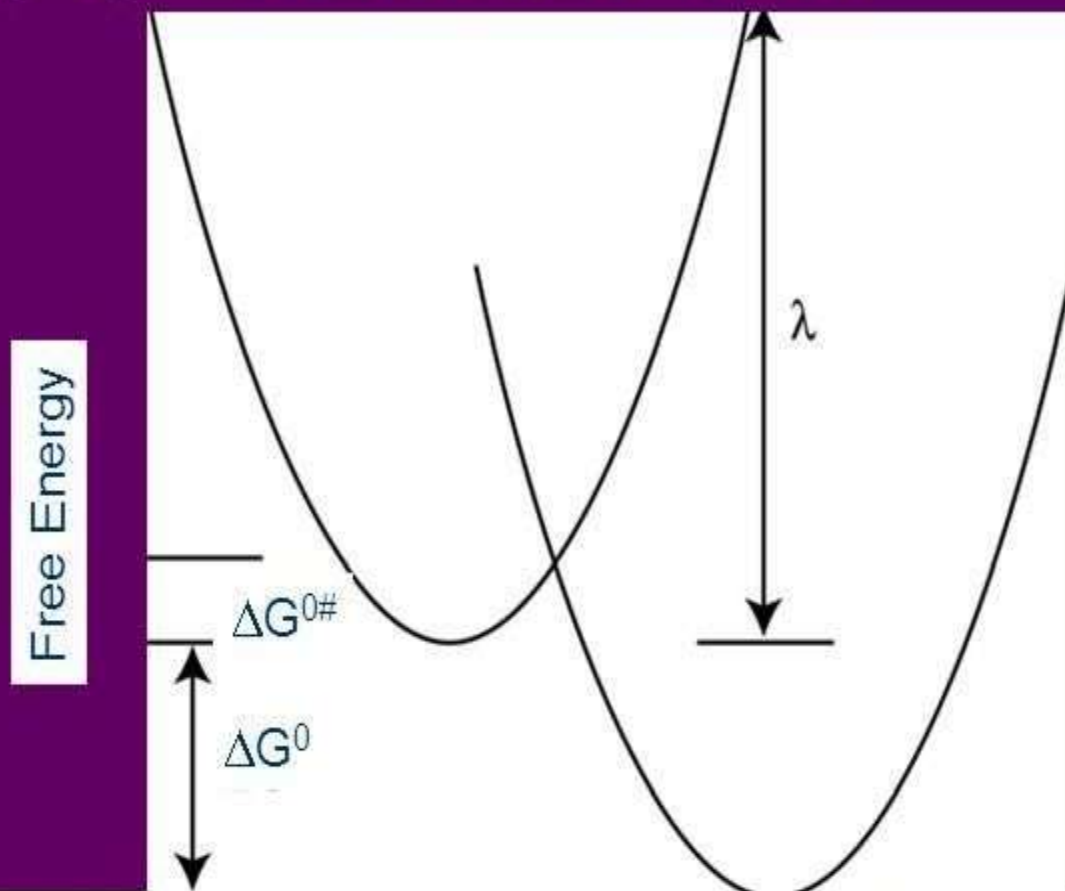
(Nobel Prize in Chemistry in 1992)



$$k_{\text{et}} = A \exp^{-\Delta G^\ddagger / kT} = (4\pi^2/h) H_{\text{RP}}^2 (4\pi\lambda kT)^{-1/2} \exp^{-(\Delta G^\circ + \lambda)^2 / 4\lambda kT}$$

k_{et} is the rate of electron transfer

- The **reorganization energy** λ is defined as the energy required to “reorganize” the structure from initial to final coordinates, without changing the electronic state.



Reaction coordinate: contains positions of atoms (nuclear coordinates)

1. For a reaction D and A must diffuse together. They form the precursor complex, usually a kinetic, unstable, solvated encounter complex.

2. which by electron transfer is transformed to the successor complex, and finally this separates by diffusion. For a one electron transfer the reaction (D and A may already carry charges).

3. The electron transfer step is slow, the separation of the successor complex is fast).

Redox reactions are preferably run in polar solvents. Donor and acceptor then have a solvent shell and the precursor and successor complexes are solvated also. The closest molecules of the solvent shell, or the ligands in complexes, are tightly bound and constitute the "inner sphere". Reactions in which these participate are called inner sphere redox reactions. The free solvent molecules constitute the "outer sphere". Outer sphere redox reactions do not change the inner sphere, neither formation of bond nor breaking of bond takes place.

References & further reading

S & Atkins, Oxford Inorganic Chemistry, Fourth Edition

Alagappa university, M. Sc. Chemistry

Pearson, Inorganic chemistry

Fmiza Hammer Synthesis & reaction of organometallic compound



Thank You

